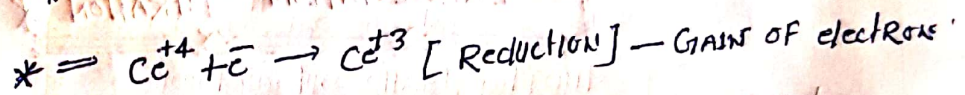
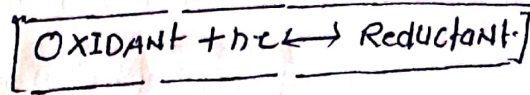


** OXIDATION - REDUCTION **

→ This TITRIMETRIC Method is mainly based upon the change of the oxidation number or electron transfer between the reactants. that is these reaction are mainly based upon the Oxidation Reduction.

→ The principle involved in the Oxidation-Reduction titration is that Oxidation process involves the loss of electron whereas the reduction process involves the gain of electron.



⊛ Oxidizing Agent ⇒ Reducing Agent

[A] → KMnO_4

[A] MIOHR SALT

[B] → $\text{K}_2\text{Cr}_2\text{O}_7$

[B] OXALIC ACID

[C] → Iodine Solution

[C] Sodium thiosulphate

INDICATOR ⇒

[A] SELF-INDICATOR ⇒ The titrant itself act as a SELF INDICATOR
It shows the INTENSE colour at the

END POINT:

Ex → Potassium Permanganate - end-point is PINK to colourless.

* → Iodine → end-point - BROWN to BLACK colour

* → CERIC AMMONIUM SULPHATE → end-point colourless to yellow colour.

EXTERNAL INDICATORS

Ex = Methylene Blue

⇒ PHENANTHROLINE BLUE ⇒ Blue-colour
*** complex

* EXTERNAL INDICATORS

Ex = FERROUS IONS