PRECIPITATION TITRATION

⇒ The Precipitation Method is Based on titration with the use of Reaction accompanied By formation of sparingly soluble compounds.

* ⇒ The Precipitate Must be Practicaly INSoluble.
* ⇒ The Precipitation Reaction Very Rabid Reaction.
* ⇒ The titration Result should Not be distorted By Adsorption effect.
* ⇒ It Must be possible to detect the equivalence point during the titration

* ⇒ Precipitation titration are mainly Based on the formation of the Precipitate By the reaction of the samble with precipitation agent.

$$Ag^+ + Cl^- \rightarrow AgCl\ (PPt)$$

*→ Requirement for the Precipitation Indicator.

⇒ The colour change should occurs Rapidly.

→ The colour change should take place with change in the titration curve.

[A] → Chromate Ion [Mohr's Method]
[B] → Fluorescein [Fajans-Method]
[C] → Iron-Ion [Volhard Method]

⇒ Factor Affecting the Precipitation titration —

- Nature of solvent.
- pH of solution
- Concentration of Reclont.
- The Nature and concentration of the foreign substance
- low solubility product formation during the titration

*→ Fluorescein ⇒ yellow green → Pink.
*→ Tartrazine ⇒ Colourless → green.
* Eosin ⇒ Pink ⇒ Reddish-Violet.