

* PRECIPITATION TITRATION *

⇒ The precipitation method is based on titration with the use of reactions accompanied by formation of sparingly soluble compounds.

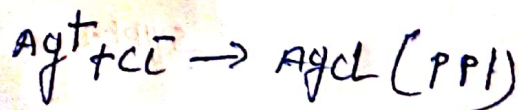
* ⇒ The precipitate must be practically insoluble.

* ⇒ The precipitation reaction very rapid reaction.

* ⇒ The titration result should not be distorted by adsorption effect.

* ⇒ It must be possible to detect the equivalence point during the titration.

* ⇒ Precipitation titration are mainly based on the formation of the precipitate by the reaction of the sample with precipitating agent.



* → Requirement for the Precipitation Indicator.

- The colour change should occur rapidly.
- The colour change should take place with change in the titration curve.

- [A] → CHROMATE ION [MOHR'S METHOD]
- [B] → FLUORESCHEIN [FAJANS-METHOD]
- [C] → IRON-ION [VOLHARD METHOD]

→ Factor Affecting the Precipitation Titration —

- Nature of solvent.
- pH of solution
- Concentration of reagent.
- The nature and concentration of the foreign substance.
- Low solubility product formation during the titration.

- * → FLUORESCHEIN ⇒ yellow green → pink.
- * → TARTRAZINE ⇒ colourless → green.
- * → EOSIN ⇒ pink = reddish-violet.

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